



**TC-TUNGSTEN** COMPOUNDS  
environment-friendly heavy liquid

# Sodium metatungstate / Sodium polytungstate \*

The ever-increasing awareness in the chemicals industry of the importance of environmental issues has resulted in an overall steady increase in the use of sodium polytungstate over recent years. We are absolutely sure that sodium polytungstate will completely replace the highly toxic, carcinogenic, chlorinated brominate hydrocarbons in the field of sink/swim analysis in future.

Sodium polytungstate – also known as sodium metatungstate – belongs to the group of so-called "non-phase-in" substances, which means that sodium polytungstate has already been registered under the requirements of the German Law on Chemical Substances and in accordance with the EC Directive 67/548/EC.

Sodium polytungstate has therefore been subjected to all the necessary statutory test procedures, a summary of the certificates for which is contained under the ELINGS No. 412-770-9.

Sodium polytungstate has automatically been allocated a registration number under the EU's REACH regulations and is therefore deemed to have satisfied the necessary registration requirements.

Sodium polytungstate has already been tested using all the currently available methods for toxicity and eco-toxicity. The present state of knowledge about the effects of sodium polytungstate on humans, animals and the environment is sufficiently high for us to be able to exclude the possibility of any risks to health arising from the proper use of this new type of heavy liquid.

This extensive series of tests has also demonstrated that sodium polytungstate may be classified as being non-toxic, which means that the use of the use of aqueous sodium polytungstate solution is set to increase even further across the whole world.

Furthermore, we should like to point out that another important benefit of aqueous sodium polytungstate is the fact that it may be reused. We will collect free of charge and for reuse any used sodium polytungstate solutions which are no longer required by customers, even if they have been diluted or contaminated.

On the following pages we shall describe the most important properties of sodium polytungstate and its aqueous solutions. We should also like to draw your attention to some interesting publications which describe its possible uses.

Sedimentation of gold leaf  
in an aqueous sodium  
polytungstate-solution  
with a density of 3,0 g/cm<sup>3</sup>



# Sodium polytungstate

SODIUM POLYTUNGSTATE is a compound, which has been successfully used in the manufacture of heavy liquids for a relatively short period. It has significant advantages when compared to the use of zinc chloride solution or the highly toxic halogenated carbons for sink/swim analysis.

Formula  $\text{Na}_6[\text{H}_2\text{W}_{12}\text{O}_{40}]$  or  $3\text{Na}_2\text{WO}_4 \cdot 9\text{WO}_3 \cdot \text{H}_2\text{O}$

Appearance White crystals or light yellow-green transparent solution

Properties Very easily soluble in water, pH-neutral solution, maximum possible density of solution:  $3,1 \text{ g/cm}^3$  at  $25^\circ \text{C}$

Features  
Non-toxic  
Non-flammable  
Odourless  
Reusable  
Density may be adjusted by dilution from  $\rho = 1,1 - 3,1 \text{ g/cm}^3$   
Low viscosity  
Easy-to-use  
Sinks and floats may be easily cleaned with water after use  
Environmentally friendly

Application The heavy liquid solution is prepared by dissolving sodium polytungstate in de-ionized water. The resulting compound is a salt which is easily soluble in water. The maximum attainable density of the solution in water at room temperature is  $3.1 \text{ g/cm}^3$ . Further chemical properties of the solution may be ascertained by studying the diagrams. These show density as a function of the sodium polytungstate content (Fig. 1) and viscosity as a function of density (Fig. 2). As may be seen from Fig. 2, viscosity increases only slightly with densities up to  $2.5 \text{ g/cm}^3$ . This process also allows the fine granules to be separated. The use of a laboratory centrifuge will speed up the separation process.



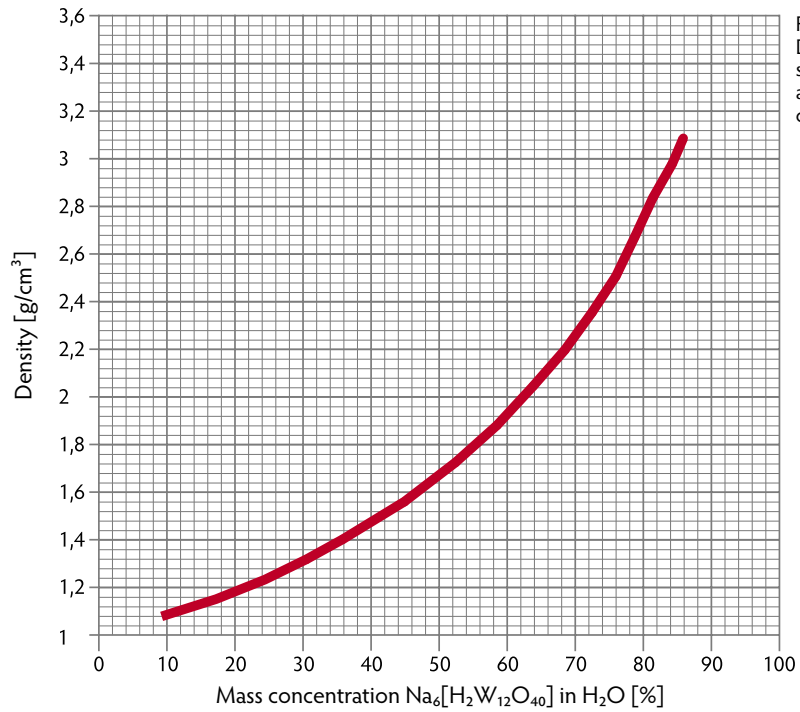


Fig. 1:  
Density of aqueous sodium  
polytungstate  
as a function of mass  
concentration

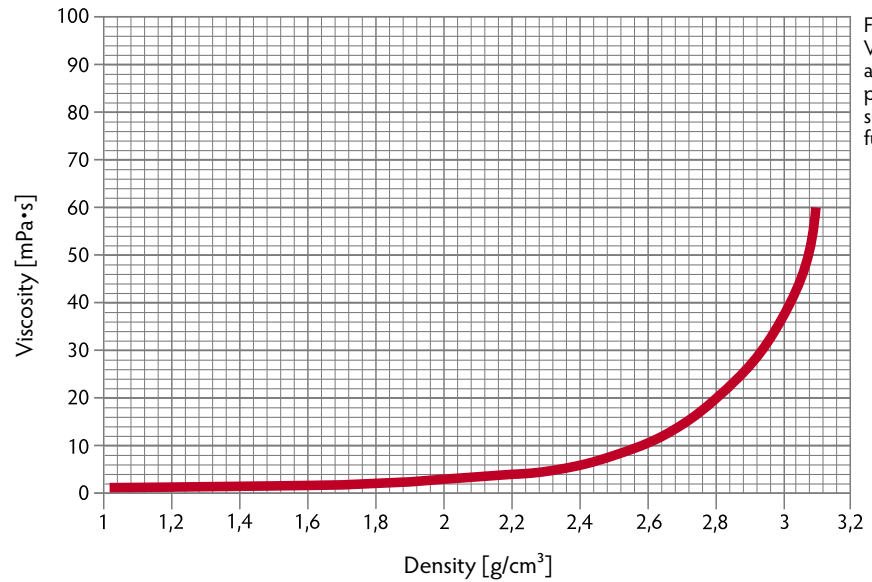



Fig. 2:  
Viscosity of aqueous sodium  
polytungstate  
solutions as a  
function of density

# Sodium polytungstate

Structur	<p>This is a 12-fold aggregated isopolytungstate with a molar mass of 2986.12 g/Mol. The structure of polytungstate is modelled using octahedrons in which the oxygen ions are located in the corners and the tungsten ions in the middle of the octahedrons. When represented as a spherical model, the oxygen ions form a dense spherical shell while the tungstate ions fill the open spaces in the octahedra.</p> <p>Because of this structure this substance may be considered a "true" metatungstate, represented structurally as <math>\text{Na}_6[\text{H}_2\text{W}_{12}\text{O}_{40}]</math>. It is known that in "true" metatungstates both of the oxygen ions are located in the central empty space of the polyanion and that they are unable to penetrate the external shell of the spherical casing.</p> <p>Solid crystalline sodium polytungstate is not hygroscopic and may be stored for an unlimited period at room temperature.</p>
Notes	<p>The following points should be observed when using aqueous polytungstate solutions:</p> <ol style="list-style-type: none"><li>Only distilled or de-ionized water should be used.</li><li>All containers must be sealed properly after use.</li><li>Only containers made from glass, synthetic materials or stainless steel should be used</li><li>Do not bring the solution into contact with reduced substances. The resultant blue colouring will not, however, affect the achievement of the required density in any way. Colour may be removed from the solution and a blue colouration prevented by the addition of a few drops of hydrogen peroxide.</li><li>The Sinks and floats used should not contain any water soluble ions. <math>\text{Pb}^{2+}</math>-, <math>\text{Ag}^+</math>-, <math>\text{Sn}^{2+}</math>- and <math>\text{Ba}^{2+}</math>- ions in particular form deposits which are difficult to dissolve. If these ions should be present in the samples these must be washed with hot water prior to the – sink/swim analysis.</li></ol>
Analysis	<p>Solid sodium polytungstate contains at least <math>86\% \pm 1\% \text{WO}_3</math>.</p> <p>The quantity of bound water in sodium polytungstate may vary slightly. Typical analysis values (not guaranteed values) are: <math>\text{Al} \leq 0.0015\%</math>; <math>\text{As} \leq 0.012\%</math>; <math>\text{Cu} \leq 0.001\%</math>; <math>\text{Fe} \leq 0.005\%</math>; <math>\text{Mo} \leq 0.005\%</math>; <math>\text{Si} \leq 0.005\%</math>; <math>\text{Ti} \leq 0.001\%</math>.</p>
Form of supply	<p>A standard delivery of sodium polytungstate will be of a ready-to-use aqueous solution with a density of <math>2.82 \pm 0.02 \text{ g/cm}^3</math> or with densities of <math>\geq 3.0 \text{ g/cm}^3</math> in 1 kg, 5 kg, 10 kg and 25 kg. Sodium polytungstate may also, however, be supplied in crystalline form or as granules batches in batches of 1 kg, 5 kg, 10 kg and 25 kg.</p>





Left: Sodium polytungstate granulate  
Right: Sodium polytungstate powder  
(Image greatly magnified)

**Toxicology**

In general, tungsten compounds are classified as being non-toxic (c.f. "Metal Toxicity in Mammals – 2", Chemical Toxicity of Metals and Matalloids by B. Venugopal and T. D. Luckey, Department of Biochemistry, University of Missouri, Columbia 1978 and Handbook on the Toxicology of Metals, Chapter 39, By L. Fridlberg, G. F. Nordberg and V. B. Vouk, Elsevier/ North Holland Biomedical Press (1979)).

In addition, there is a complete list of toxic heavy metals in the Manual of Pharmacology, Lehrbuch der Pharmakologie by Bader, and this list does not include tungsten.

Sodium polytungstate / Sodium metatungstate was classified as being a new substance for the purposes of the new German Law on the use of Chemicals (Chemikaliengesetz (ChemG)) and has been fully tested in accordance with the relevant legal requirements (in relation to notification and registration obligations and evidence of testing in accordance with the Statutory Ordinance on the Evidence of Testing of Chemical Substances Prüfnachweisverordnung - ChemPrüfV).

The following toxicity test values have been ascertained for sodium polytungstate :  
LD<sub>50</sub> oral, Rat = 1715 mg/kg; LD<sub>50</sub> dermal, Rat = > 2000 mg/kg

Sodium polytungstate is therefore classified as being non-toxic for the purposes of the German Chemicals Law. Sodium polytungstate does not irritate the skin or cause sensitization.

It should be kept well away from the eyes.

**Further Details can be gathered from the safety data sheet.**

**Disposal**

Used, discarded, diluted or contaminated samples of aqueous sodium polytungstate solution will be collected by us free of charge. We are then able to recover the remaining tungsten from the solution using a specialised process.

\*) former patent by TC-Tungsten Compounds, Dr. Rainer Kamps

The data contained in this information sheet is provided free of charge and is based on technical data which TC-Tungsten Compounds considers to be reliable. This information is produced for the use by people who have the appropriate technical knowledge and may be used by such people who see themselves fit to do so and at their own risk. Because the operating condition under which these products may be used are outside of the area of our control we are unable to provide any explicit or implied guarantees or accept any form of liability in relation to the ways in which this information may be used.



# Our Products



- SPT 0** Crystalline sodium polytungstate, particularly low in carbon and nitrogen
- SPT 1** Sodium polytungstate in powder form
- SPT 2** Sodium polytungstate granulate (Lower dust component)
- SPT 3** Sodium polytungstate, ready-to-use liquid, available with various fluid densities up to  $3.0 \text{ g/cm}^3$  (Standard fluid density  $2.82 \text{ g/cm}^3 \mp 0.02 \text{ g/cm}^3$ )
- SPT 4** Sodium polytungstate / tungsten carbide suspension with densities of  $3.2 \text{ g/cm}^3$  to  $4.4 \text{ g/cm}^3$
- SPT 5** Sodium polytungstate, ready-to-use liquid, available in various fluid densities up to  $3.0 \text{ g/cm}^3$  (Standard fluid density:  $2.82 \text{ g/cm}^3 \mp 0.02 \text{ g/cm}^3$ ) Highly stabilised for use with metallic samples.





# Publications relating to the uses and applications of Sodium polytungstate

## **„Sodium metatungstate, a new medium for binary and ternary density gradient centrifugation”**

Bodo Plewinsky, Rainer Kamps;  
Makromol. Chem., **185**, 1984, 1429-1439

## **„A non-toxic heavy liquid and inexpensive filters for separation of mineral grains”**

John Callahan; J. of Sediment. Petrol., **57** (4), 1987, 765-766

## **„A nontoxic substitute for hazardous heavy liquids-aqueous sodium polytungstate (3Na<sub>2</sub>WO<sub>4</sub>·9WO<sub>3</sub>·H<sub>2</sub>O) solution”**

Murray R. Gregory, Keith A. Johnston;  
J. of Geol. and Geophys., **30**, 1987, 317-320

## **„The use of sodium polytungstate for conodont separations”**

Norman M. Savage; J. Micropalaeontol., **7** (1), 1988, 39-40

## **„Sodium metatungstate: a new heavy-mineral separation medium for the extraction of conodonts from insoluble residues”**

Stanley T. Krukowski;  
J. Paleont., **62** (2), 1988, 314-316

## **„The use of sodium polytungstate in heavy mineral separation”**

Michael Torresan, United States Department of the Interior (Geological Survey (US)),  
Open-File Report **87-590**, 1987

## **„Recycling of Sodium polytungstate used in soil organic matter studies”**

J. Six, P. A. Schultz, J.D. Jastrow, R. Merckx;  
Soil Biol. & Biochem., **31**, 1999, 1193-1196

## **„Sodium metatungstate as a medium for measuring particle density using isopycnic density gradient ultracentrifugation”**

Mark D. Hoover, Gregory L. Finch, T. Castorina;  
J. Aerosol Sci., **22** (2), 1991, 215-221

## **„The use of sodium polytungstate for the separation and concentration of living dinoflagellate cysts from marine sediments”**

Bolch, C.J.S.; Phycologia, **36** (6), 1997, 472-478

## **„Density Separations in Heavy Inorganic Liquid Suspensions”**

D. Rhodes, S. T. Hall and N. J. Miles;  
XVIII International Mineral Processing Congress, Sydney, 23-28 May 1993

## **„Improved density gradient separation techniques using Sodium Polytungstate and a comparison to the use of other heavy liquids”**

Gary L. Skipp & Isabelle Brownfield; U.S. Department of the Interior U.S. (Geological Survey), Open-File Report **92-386**

## **„Laboratory procedures for processing tephra samples”**

DeAnne S. Pinney, Alaska Division of Geological and Geophysical Surveys,  
Public-Data File **91-30**, 1991

## **„Thermoluminescence of foods: Origins and implications for detecting irradiation”**

D. C. W. Sanderson, C. Slater and K. J. Cairns;  
Radiat. Phys. Chem., **34** (6), 1989, 915-924

## **„Detection by Thermoluminescence of an Irradiation Treatment of Five Species of Dehydrated Fruit and Vegetables Report on a CTCPA/AIFLD International Interlaboratory Study Edited ”**

Eric Marchioni and Henry Delincée;  
Berichte der Bundesforschungsanstalt für Ernährung, BFE-R-99-02

## **„Detection of irradiated components in flavour blends composed of non-irradiated spices, herbs and vegetable seasonings by thermoluminescence method”**

K. Malec-Czechowska, W. Stachowicz;  
Nukleonika, **48** (3), 2003, 127-132

## **„Thermoluminescence Detection of Irradiated Fruit Vegetables: International Interlaboratory Trial”**

David C. W. Sanderson, Lorna A. Carmichael, Saffron Fisk; J. of AOAC, **86** (5), 2003, 971-975

## **„Thermoluminescence (TL) of Minerals Separated from Irradiated Mussel”**

Sang-Duk Yi and Man-Jin Oh;  
J. Food Sci. Nutr., **10**, 2005, 17-21

## **„Upper Pleistocene deposits of the Comprida Island (São Paulo State) dated by thermoluminescence method”**

Kenitiro Suguio, Sonia H. Tatum, Emilia A. Kowata, Casimiro S. Munita, Rosemeire P. Paiva;  
An. Acad. Bras. Ciênc., **75** (1), 2003

## **„Aerobactin production by a planktonic marine Vibrio sp.”**

S.M.C. Robinson, R. A. Chandler;  
Limnol & Oceanogr., **38** (5), 1993, 1091-1097

## **„An effective and safe method for sorting small molluscs from sediment”**

Robinson, S. M. C.; Chandler, R. A.;  
Limnol. & Oceanogr., **38** (5), 1993, 1088-1091

## **„Study of Free and Occluded Particulate Organic Matter in Soils by Solid-state <sup>13</sup>C CP/MAS NMR Spectroscopy and Scanning Electron Microscopy”**

A. Golchin, J. M. Oades, J. O. Skjemstad and P. Clarke; Aust. J. Soil Res., **32**, 1994, 285-309

## **„Miocene to quaternary paleoceanography in the Northern North Atlantic: Variability in carbonate and biogenic opal accumulation”**


G. Bohrmann, R. Henrich and J. Thiede, Geological History of the Polar Oceans: Arctic Versus Antarctic, Kluwer Academic Publishers, 1990, 647-675

## **„Separation and identification of the silt-sized heavy-mineral fraction in sediments”**

Commeau, J. A.; Poppe, L. J.; Commeau, R. F.;  
U.S. Geological Survey Circular, **1071**, 1992, 13

## **„Soil Structure and Carbon Cycling”**

A. Golchin, J. M. Oades, J. O. Skjemstad and P. Clarke; Aust. J. Soil Res., **32**, 1994, 1043-1068



Copper/aluminium  
separation process  
in a highly stabilized  
sodium polytungstate solution

**„Incorporation of nitrogen from decomposing red alder leaves into plants and soil of a recent clearcut in Oregon”**

Swanston, C. W. Myrold, D.;  
Canad. J. Forest Res., **27** (9), 1997, 1496-1502

**„Compartmentalization and Coupling of Nitrogen Cycling in Soil”**

David D. Myrold, Peter J. Bottomley;  
Award Number: 97-35107-4357, 1999-2000

**„Using lateral capillary forces to compute by self assembly”**

Paul W. K. Rothemund;  
PNAS, **97** (3), 2000, 984-989

**„Serrabrancaite, MnPO<sub>4</sub>·H<sub>2</sub>O, a new mineral from the Alto Serra Branca pegmatite, Pedra Lavrada, Paraiba, Brazil”**

Thomas Witzke, Reinhard Wegner, Thomas Doring, Herbert Pöllmann and Walter Schuckmann; American Mineralogist, **85**, 2000, 847-849

**„Testing optically stimulated luminescence dating of sand-sized quartz and feldspar from fluvial deposits”**

J. Wallinga, A. S. Murray, G.A.T. Duller, T. E. Törnqvist; Earth & Planet. Sci. Let., **193**, 2001, 617-630

**„Preferable use of red-thermoluminescence (RTL)-dating for quartz extracts from archaeologically burnt pottery –comparison of RTL and BTL (blue-TL) measurements using single-aliquot regenerative-dose (SAR) method”**

Tetsuo Hashimoto, Takashi Yawata, and Masato Takano; Ancient TL, **21** (2), 2003

**„An attempt to separate Q from the Allende meteorite by physical methods”**

Sachiko Amari, Shiho Zaizen, and Jun-Ichi Matsuda; Geochim. et Cosmochim. Acta, **67** (24), 2003, 4665-4677

**„Oxidation of styrene and cyclohexene under microwave conditions”**

Janine Freitag, Matthias Nüchter and Bernd Ondruschka; Green Chemistry, **5**, 2003, 291-295

**„Präzise Dichtemessungen mit schwebenden Festkörpern”**

Hans Toth, Heinz Fehlauer;  
Phys. Unserer Zeit, **35** (2), 2004

**„Sedimentation and carbon burial on the northern California continental shelf: the signatures of land-use change”**

Elana L. Leithold, David W. Perkey, Neal E. Blair, Todd N. Creamer;  
Continental Shelf Research, **25**, 2005, 349-371

**„Alternative saline solutions to float foraminiferal tests”**

D. L. Semensatto, Jr. and D. Dias-Brito;  
J. of Foraminiferal Res., **37** (3), 2007, 265-269



**TC-TUNGSTEN** COMPOUNDS  
environment-friendly heavy liquid

TC-Tungsten Compounds  
Dr. Rainer Kamps  
Tannenweg 2  
D-96271 Grub a. Forst  
Germany  
Tel.+49 (0) 9560 742  
Fax +49 (0) 9560 1707  
[www.heavy-liquid.com](http://www.heavy-liquid.com)